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Ans. Acetic acid dissociates to produce one H+(aq) ion per molecule and hence its basicity is one. $\text{CH}_3\text{COOH} + \text{H}_2\text{O} \rightleftharpoons \text{CH}_3\text{COO}^- + \text{H}^+$ Magnesium hydroxide reacts completely with $2\text{H}^+(\text{aq})$ ions of an acid to form salt and water, therefore its acidity is 2. $\text{Mg}(\text{OH})_2 + \text{H}_2\text{SO}_4 \rightarrow \text{MgSO}_4 + 2\text{H}_2\text{O}$ 11. Why is sulphuric acid a dibasic acid? Give three reasons.

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from the molecule of an acid, with metal ions, is called an acid salt. For example, 1. Sodium hydrogen sulphate (NaHSO_4) 2. Calcium hydrogen carbonate [$\text{Ca}(\text{HCO}_3)_2$]. (iii) Basic salt — A salt formed by the partial neutralisation of hydroxyl ions (OH^-) of a base, by an acid is called a basic salt. For example,

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while acetic acid (CH_3COOH) is added in test tube B. In which test tube will the fizzing occur more vigorously, and why? Ans. Fizzing will occur more vigorously in test tube A as compared to the test tube B. It is because, the concentration of $\text{H}^+(\text{aq})$ ions in hydrochloric acid, which is a strong acid, is more than acetic acid, which is a weak ...

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Q.10 Equal lengths of magnesium ribbons are taken in test tubes A and B. Hydrochloric acid (HCl) is added to test tube A, while acetic acid (CH_3COOH) is added to test tube B. In which test tube will the fizzing occur more vigorously and why? Sol. In test tube A, fizzing occur more vigorously because HCl is a strong acid and dissociate more.

Acids Bases and Salts : NCERT Exercise Questions ...

Acids, Bases and Salts Physical and Chemical Changes Weather, Climate And Adaptation Wind Storms and Cyclones Soil Respiration in Organisms Transportation in Animals and Plants Reproduction in Plants Motion and Time Electric Current and its Effect Light Water A Precious Resource Forests: Our Lifeline Waste Water Story

Acids, Bases and Salts CBSE Science Class 7 Chapter Wise ...

Acid-Base Reactions Reactions of acids and bases a) Reaction of acids and bases with metals. Acid + active metal → salt + hydrogen + heat. $2\text{HCl} + \text{Mg} \rightarrow \text{MgCl}_2 + \text{H}_2$ (†) Base + metal → salt + hydrogen + heat. $2\text{NaOH} + \text{Zn} \rightarrow \text{Na}_2\text{ZnO}_2 + \text{H}_2$ (†) A more reactive metal displaces the less reactive metal from its base. $2\text{Na} + \text{Mg}(\text{OH})_2 \rightarrow 2\text{NaOH} + \text{Mg}$

Class 10 Chapter 2 Science Notes Acids, Bases, and Salts

Chemical properties of acids: i) Acids react with active metals to give hydrogen gas. $\text{Zn} + \text{H}_2\text{SO}_4 \rightarrow \text{ZnSO}_4 + \text{H}_2$. ii) Acids react with metal carbonate and metal hydrogen carbonate to give carbon dioxide. $\text{NaHCO}_3 + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O} + \text{CO}_2$. iii) Acids react with bases to give salt and water. This reaction is called as neutralization reaction. $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$

acids bases and salts for class 10 cbse notes

Acids, Bases, and Salts 31,985 Many acids and bases occur naturally in nature, such as citric acid in fruits like orange, lemon etc., tartaric acid in tamarind, malic acid in apples and lactic acid in milk and milk products, hydrochloric acid in gastric juices. Similarly, many bases are found such as lime water.

Acids, Bases, and Salts - Introduction, Dissociation ...

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Reaction of Acids with Metal Carbonates and Metal Hydrogencarbonates. • Acids reacts with Metal Carbonates and Metal Hydrogencarbonates to form Salt, Carbon dioxide and water. Metal carbonate/Metal hydrogen carbonate + Acid → Salt + Carbon dioxide + Water. • Examples: (i) $2\text{HCl} + \text{Na}_2\text{CO}_3 \rightarrow 2\text{NaCl} + \text{CO}_2 + \text{H}_2\text{O}$.

Notes of Ch 2 Acids, Bases and Salts | Class 10th Science

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Online Test of Chapter 2 Acid, Bases and Salt MCO 1. Science| Class 10th Questions: 1. Basic characteristics of Acid is/are? i) Sour taste ii) Turns blue litmus red iii) Turns red litmus blue iv) Both (i) and (ii) 2. Basic Characteristics of Base is/are? i) Sour taste ii) Turns red litmus blue iii) Turns blue litmus red 3. Which acid do we...

Chapter - 2 Acid, Bases and Salt MCO Test 1| Class 10th ...

Acids, bases and salts affect chemistry as well as our day to day life. They can be easily identified by their taste; that is acids taste sour and bases taste bitter and salts itself have salty taste. Acids are usually found in many substances including various food items but their presence in many fruits is very prominent, for example:

Acids Bases and Salts | Properties of Acids, Bases and Salts

Acids give citrus fruit its sour taste, while bases such as ammonia are found in many types of cleaners. Salts are a product of the reaction between an acid and a base. A common method used to determine an acid or a base is a litmus test, but there are other characteristics that can help you identify acids, bases and salts.

Characteristics of Acids, Bases & Salts | Sciencing

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Acids and bases interact with each other in what is called a neutralization reaction. The products of the reaction are a salt and water. The products of the reaction are a salt and water. The pH is "neutralized" to 7 if both the acid and base fully react.

Acids and Bases Chemistry Quiz - ThoughtCo

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Acids, Bases & Salts | Class 10 Chemistry | Science Chapter 2

Class 7 Chapter 5 - Acids, Bases and Salts Neutralization Neutralization is a chemical reaction in which acid and base react to form salt and water. Hydrogen (H+) ions and hydroxide (OH-) ions ...

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